

## Sorption Reversibility of Cadmium from Aqueous Solutions on Natural Firoozkoh Zeolite

F Pooladi<sup>1</sup> and M Hamidrpour<sup>1\*</sup>

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### Abstract

Removal of boron from aqueous environments (soil and water) is difficult, because it is present as  $B(OH)_3$  and  $B(OH)_4^-$  species. This research was done to study the sorption of boron by HDTMA-modified zeolite. The sorption of B on modified zeolite was studied as a function of pH (B concentration: 1 and 10  $mg L^{-1}$ ) in the range of 6-9.5, and as a function of ionic strength (0.03 and 0.06 M  $Ca(NO_3)_2$  or  $Mg(NO_3)_2$ ) at a constant B concentration of 5  $mg L^{-1}$ . Sorption isotherm was performed for the solutions containing initial B concentration in the range of 1-15  $mg L^{-1}$  using a 24h batch equilibration experiment. The results revealed that surfactant-modified zeolite exhibited the best performance at pH 9.5, and sorption of B increased with the increase of suspension pH. Greater B adsorption in the Ca system over the Mg system was clearly observed for the modified zeolite. Sorption isotherm of B were well described by the Freundlich and Langmuir models but the Freundlich sorption model described the interaction between B and the mineral material better than the Langmuir model. Maximum sorption capacity ( $q_{max}$ ) of the sorbent was 120  $mmol kg^{-1}$ . The experimental data showed that HDTMA-modified zeolite used in this study had a reasonable sorption capacity for B.

**Keywords:** Boron, Sorption isotherm, Modified zeolite, Removal.

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1. Dept. of Soil Sci., Isf. Univ. of Technol., Isfahan, Iran.

2. Dept. of Soil Sci., Vali-e-Asr University of Rafsanjan, Rafsanjan, Iran.

\*: Corresponding Author, Email: mohsen\_hamidpour@yahoo.com